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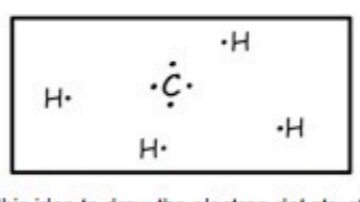
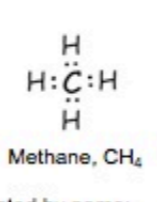
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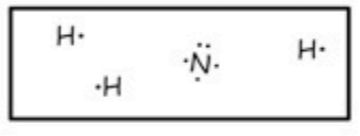
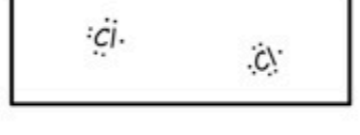
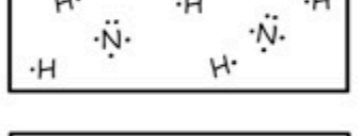
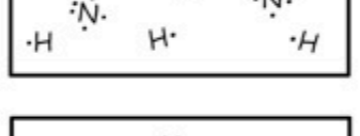
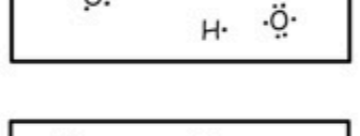
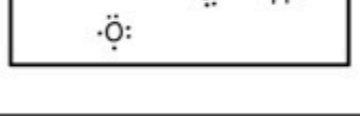
Conceptual Chemistry

How Atoms Bond
Electron-Dot Structures

Electron-dot structures tell us how atoms tend to bond with other atoms. Carbon's electron-dot structure, for example, shows 4 unpaired valence electrons. Each hydrogen atom has 1 unpaired electron. Unpaired electrons from different atoms can pair up, resulting in a bond. For carbon and hydrogen, we have the following, which creates the molecule methane, CH₄:

	→	
		Methane, CH ₄

Use this idea to draw the electron-dot structures of the products indicated by name:

	→	Ammonia, NH ₃
	→	Chlorine, Cl ₂
	→	Hydrazine, N ₂ H ₄
	→	Nitrogen, N ₂ + Hydrogen, H ₂
	→	Hydrogen peroxide, H ₂ O ₂
	→	Oxygen, O ₂ + Hydrogen, H ₂

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Chemistry (12th Edition) answers to Chapter 6 - The Periodic Table - 6.3 Periodic Trends - 6.3 Lesson Check - Page 182 18 including work step by step ... Section 6.3 Periodic Trends 171 Group Trends in Atomic Size In the Figure 6.14 graph, atomic radius is plotted versus atomic number. Look at the data for the Quiz Worksheet - Atomic and Ionic Radii on the Periodic Table the atomic radius is ... Chapter 6.3 6.4 Atomic/Ionic Radius Worksheet Name: _____ Chemistry 2 Introduction To The Periodic Table Worksheet Answers E. Modern Periodic Table 1. Modern Periodic Law: similar properties of the elements occur periodically when ...

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